



# SHIKSHA CLASSES

Subject : Science -I

## Question Paper

Total Marks : 20

Class : X

### 9. Carbon Compounds

Time : 1 Hour

**Q. 1 : A) Choose the correct alternative :**

2

1) The first compound to be prepared in the laboratory was -----.

- a) Methane                      b) Ethyl alcohol  
c) Acetic acid                  d) Urea.

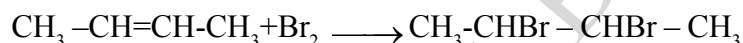
2) The number of C-H bonds in ethane  $C_2H_6$  molecule are -----.

- a) 4                                  b) 6  
c) 8                                  d) 10

**B) Solve the following questions. (Any One)**

1

1) Identify the type of reaction :



2) Write the structural formula for :

1 – bromopropane.

3) Find the odd one out :

Polysaccharide, Teflon, Polyvinylchloride, PET.

**Q. 2 : A) Give scientific reason. (Any One)**

2

- 1) Iodine test is used for detection of a multiple bond in a carbon compound.  
2) In hydrogenation of vegetable oil nickel catalyst is used.

**B) Solve the following questions. (Any Two)**

4

- 1) What cause the existence of very large number of carbon compound?  
2) Define structural isomerism.  
3) Saturated hydrocarbons are classified into three types. Write these names by giving one example each.  
4) Draw the structural formula for the following IUPAC name.  
(i) 2 bromo-butene      (ii) Butanoic acid

**Q. 3 : Solve the following questions. (Any Two)**

6

- 1) Write the difference between saturated hydrocarbon and unsaturated hydrocarbon.  
2) What is combustion of carbon compounds? Explain with example.  
3) What is meant by vinegar and gasohol? What are their uses?  
4) Write the IUPAC names of the following structural formulae.

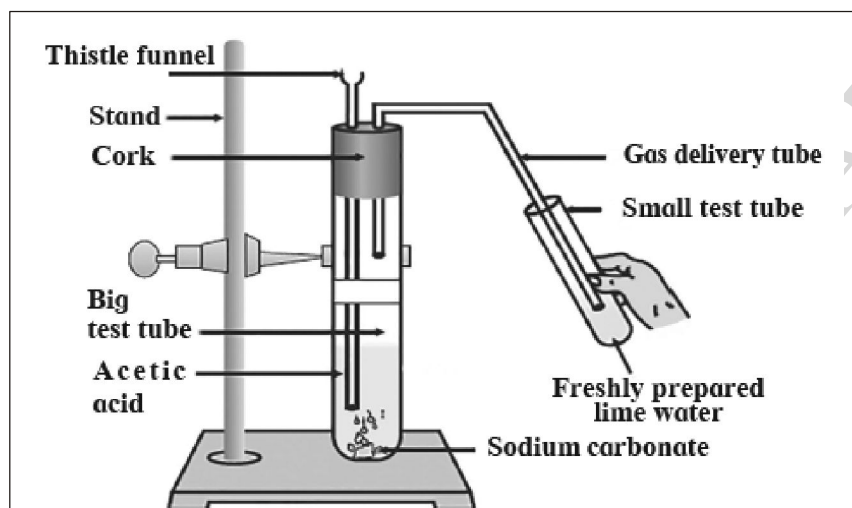
- (i)  $\text{CH}_3 - \underset{\text{Br}}{\text{CH}} - \text{CH}_3$
- (ii)  $\text{CH}_3 - \text{CO} - \text{CH}_2 - \text{CH}_3$
- (iii)  $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{OH}$

**Q. 4 : Solve the following questions. (Any One)**

5

1) Explain chemical properties of ethanoic acid.

2) With reference to given diagram answer the following question.



- Which gas does come out as effervescence in the big test tube?
- Which reaction shown in the given diagram?
- Why are bubbles seen in the small test tube?
- What is the colour change in the lime water? Write the related equation.
- Write the balanced chemical reaction of acetic acid and sodium carbonate.

\* \* \*

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