



**Q.7 :** Read the following and answer any two question from 5(i) to 5(iii). (2)

An ore on treatment with dilute hydrochloric acid produces brisk effervescences.

- i) What type of ore is this?  
a) Bauxite            b) Hematite            c) Cinnabar            d) Carbonate
- ii) By which process the metal is obtained?  
a) Roasting            b) Calcination  
c) Both a and b            d) None of the above
- iii) This type of ore is heated in.  
a) Limited supply of air            b) Complete supply of air  
c) In absence of air            d) None of these

Q.8 : Which one of the following properties is not generally exhibited by ionic compounds?

- a) Solubility in water            b) Electrical conductivity in solid state  
c) High melting and boiling points            d) Electrical conductivity in molten state

Q.9: Which of the following non-metal is lustrous?

- a) Sulphur            b) Oxygen            c) Nitrogen            d) Iodine

Q.10: An element X is soft and can be cut with a knife. This is very reactive to air and cannot be kept open in air. It reacts vigorously with water. Identify the element from the following

- a) Mg            b) Na            c) P            d) Ca

Q.11: Reaction between X and Y forms compound Z. X loses electron and Y gains electron. Which of the following properties is not shown by Z?

- a) Has high melting point            b) Has low melting point  
c) Conducts electricity in molten state            d) Occurs as solid

Q.12: The ability of metals to be drawn into thin wires is known as

- a) ductility            b) malleability            c) sonority            d) conductivity

Q.13: The electronic configurations of three elements X, Y and Z are X - 2, 8; Y - 2, 8, 7 and Z - 2, 8, 2. Which of the following is correct?

- a) X is a metal.            b) Y is a metal.  
c) Z is a non-metal.            d) Y is a non-metal and Z is a metal.

Q.14 : The atomic numbers of four elements A, B, C and D are 6, 8, 10 and 12 respectively. The two elements which can react to form ionic bonds (or ionic compound) are:

- a) A and D            b) B and C            c) A and C            d) B and D

**SECTION (B)**

**(Each - 2 Mark)**

Q.15 : Define amphoteric oxides.

**OR**

Define corrosion

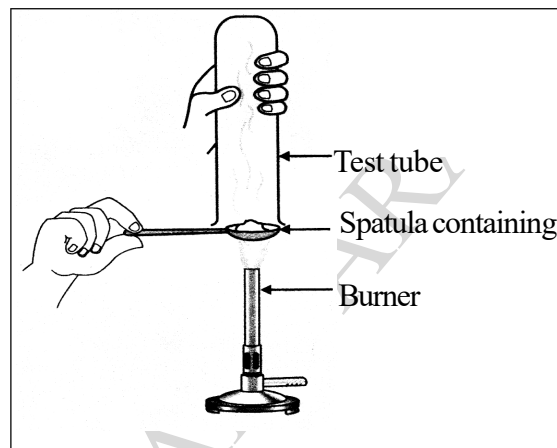
Q.16 :What is meant by aqua regia? Which metal is dissolved in aqua regia?

**SECTION (C)**

**(Each - 3 Mark)**

Q.17 :Pratyush took sulphur powder on a spatula and heated it. He collected the gas evolved by inverting a test tube over it, as shown in fig.

- a) What will be the action of gas on
- Dry litmus paper?
  - Moist litmus paper?
- b) Write a balanced chemical equation for the reaction taking place.



**Collection of Gas**

**OR**

- : Name two metals which react violently with cold water. Write any observation you would make when such a metal is dropped into water. How would you identify the gas evolved, if any, during the reactions?

Q.18 : What is cinnabar? How is metal extracted from cinnabar? Explain briefly.

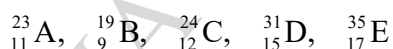
**SECTION (D)**

**(5 Mark)**

Q.19 :Metal M is the most abundant metal in the Earth crust. Its electronic configuration is 2, 8, 3 and occurs in the Earth crust as its oxides  $M_2O_3$ . Name the ore of the metal containing  $M_2O_3$ . Explain the method used for concentration of the ore. Also give reason for choosing that method.

**OR**

- :i) How do you classify elements into metals and non-metals on the basis of their electronic configuration? Choose metal and non-metal out of the following :



- ii) What type of bond will be formed if
- 'A' combines with 'B'
  - 'A' combines with 'E'
  - 'C' combines with 'E'
  - 'D' combines with 'E'

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