

Q.6) Give an example of polyatomic ion.

OR

Define law of conservation of mass.

Q.7) **Read the following paragraph and answer any two question from 5(i) to 5(iii)** (2)

Raunak took 5 moles of carbon atoms in a container and Krish also took 5 moles of sodium atoms in another container of same weight.

- i) Whose container is heavier?
a) Raunak b) Krish c) Both a and b d) None of them
- ii) Whose container has more number of atoms?
a) Raunak b) Krish c) Both a and b d) None of them
- iii) Mass of sodium atoms carried by Krish is
a) 115g b) 60g c) 115.5g d) 60g

Q.8) Which of the following statements is not true about an atom?

- a) Atoms are not able to exist independently.
b) Atoms are the basic units from which molecules and ions are formed.
c) Atoms are always neutral in nature.
d) Atoms aggregate in large numbers to form the matter that we can see, feel or touch.

Q.9) 1 u or 1 amu means

- a) 1/12th mass of C-12 atoms b) Mass of C-12 atom
c) Mass of O-16 atom d) Mass of Hydrogen molecule

Q.10) Laws which explain the formation of many oxides by nitrogen is

- a) Law of conservation of mass b) Law of multiple proportions
c) Law of definite proportions d) Avogadro's law

Q.11) What would be the gram molecular mass of solid sulphur?

- a) 256 g b) 128 g c) 64 g d) 32 g

Q.12) How many moles are present in 560 g of iron?

- a) 1 mole b) 10 moles c) 100 moles d) 0.1 mole

Q.13) According to the law of definite proportions, _____.

- a) matter remains constant
b) matter can neither be created nor destroyed matter remains constant
c) a chemical compound is always made up of the same element combined together in the same fixed proportion matter remains constant
d) All of these

Q.14) When alpha-particles are sent through a thin metal foil, most of them go straight through the foil because, _____.

- a) alpha-particles are much smaller than electrons.
b) alpha-particles are positively charged
c) most part of the atom is empty space
d) alpha-particles move with a low velocity

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